Ionic Bonding Worksheet #1

- 1) What are the electrons in the outer shell of an atom that form the chemical bonds called?
- 2) What are the three main types of bonds?
- 3) Using the terms: metal and nonmetal, describe what types of atoms make up each of the three types of bonds mentioned above.
- 4) What is the name of the type of chemical bond that is formed when two electrons are shared?
- 5) What is different between an ionic bond and the bond in the above question?
- 6) How many valance electrons are there in carbon? Where do you look on the periodic table to find out?

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9) 10) 11)	b . How ma How ma Explain Explain	ny Nit how to how to	d bon atoms are in CC rogen atoms are in o name binary ionic o name polyatomic io	•	f . gen atoms a			g . Sulfur at		
		ollowin	g ionic compounds:	16) Ch/NO)	10) 1400		20) Ph/NO	`		
	LiOH Na ₂ 50 ₄		14) SbCl ₃ 15) Al(OH) ₃	16) Sb(NO ₃) ₃ 17) Al ₂ (SO ₄) ₃			20) Pb(NO ₃ 21) K ₂ SO ₃	3) 2		
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/)	Label Tr	ne tollo a . N	-	valent, compounds/		d . HB)n		f. Ale)n
		b . 6		c. K₃P		e . CC			1. A16 9 . H ₂ (
 8) How many Carbon atoms are in CO₂? How many oxygen atoms are in CO₂? 9) How many Nitrogen atoms are in (NH₄)₂S? How many Hydrogen atoms? How many Sulfur atoms? 10) Explain how to name binary ionic compounds 11) Explain how to name polyatomic ionic compounds Name the following ionic compounds: 										
			-	16) Sb(NO ₃) ₃	1	8) HgO)	20) Pb(N	1O ₃) ₂	
13)	Na ₂ 50 ₄	+	15) Al(OH)₃	17) Al ₂ (SO ₄) ₃	1	9) Fe ₂ S	53	21) K ₂ SO	D ₃	
 7) Label the following as ionic, or covalent, compounds/molecules a. NaCl b. CO₂ c. K₃P d. HBr d. AlBr₃ e. CCl₄ g. H₂O 8) How many Carbon atoms are in CO₂? How many oxygen atoms are in CO₂? 9) How many Nitrogen atoms are in (NH₄)₂S? How many Hydrogen atoms? How many Sulfur atoms? 10) Explain how to name binary ionic compounds 11) Explain how to name polyatomic ionic compounds 										
		ollowin	g ionic compounds:					_		
-	LiOH		14) SbCl ₃	16) Sb(NO ₃) ₃	18) HgO		20) Pb(NO ₃	3)2		
7) 8) 9) 10) 11)	How ma How ma Explain Explain	ne follo a. N b. C ny Car ny Nit how to how to	NaCl :O ₂ bon atoms are in CC	•	gen atoms a		l ₄ D ₂ ?	!	f . AlB g . H₂(oms?	-
-	LiOH		14) SbCl ₃	16) Sb(NC) ₃) ₃	18)	HgO			20) Pb(NO ₃) ₂
13)	Na ₂ 50 ₄	+	15) Al(OH) ₃	17) Al ₂ (50) ₄) ₃	19)	Fe ₂ S ₃			21) K ₂ SO ₃

7) Label the following as ionic, or covalent, compounds/molecules